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particle with an initial kinetic energy K heading directly for the nucleus of a gold atom (charge = $+79e$).

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Table. **** The idea that all matter is
composed of atoms is fundamental to
our modern view of the world.

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The stimulated emission is in the same direction as the incident photon, and the

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Atomic Physics

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experiments show that the mass of the nucleus is less than the above. That is, mass $(Zm_p + Nm_n)$ We shall designate the difference between the computed mass in Equation 31.1 and the real mass as the mass defect. $\Delta m =$

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$Zm_p + Nm_n - M$ (31.3) where M

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experiments show that the mass of the nucleus is less than the above. That is, $mass < (Zm_p + Nm_n)$ We shall designate the difference between the

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computed mass in Equation 31.1 and the real mass as the mass defect. $\Delta m = Zm_p + Nm_n - M$ (31.3) where M is the measured mass of the nucleus. We shall now make use of Albert

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Physics pioneers Balmer, Rydberg, and Ritz discovered in atomic spectra a) that electrons behave as standing waves b) mathematical order c) a randomness that opened the door to the nature of

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probabilities in atomic reality d) that all atoms are essentially the same size e) that electrons occupy well-defined shells about the atomic nucleus

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Chapters 31 Atomic Physics -

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University of Houston

Chapter 31. Nuclear Physics. In this chapter the following topics will be discussed: The Cause of Radioactivity; Isotopes of an Element; Types of Radiations; Radioactive Decay; Carbon Dating; Nuclear Energy; How Do Reactors Work? 1. The Cause of Radioactivity: Protons and neutrons form

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the nucleus of atoms.

Chapter 31

CHAPTER 31 NUCLEAR PHYSICS AND RADIOACTIVITY ANSWERS TO FOCUS ON CONCEPTS QUESTIONS 1. (e) In the notation the symbol Z represents the number of protons, and the symbol A represents the number of protons plus

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the number of neutrons. In, $A - Z = 2$. (c) Because the atomic number of the nucleus is 37, there are 37 protons in the nucleus.

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chapter 31 nuclear physics and

radioactivity conceptual questions _____

1. REASONING AND SOLUTION Isotopes are nuclei that contain the same number of protons, but a different number of neutrons. A material is known to be an isotope of lead, although the particular isotope is not known.

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